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University koretsm@enr.orst.edu. 26.1

(a) The Clausius-Clapeyron equation:
$$\ln \left(\frac{P_2}{P_1} \right) = \frac{h_{\text{vap}}}{R} \left(\frac{1}{T_1} - \frac{1}{T_2} \right)$$

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373 [K]* +, -./ so $P_2 = 101 \text{ kPa} \left(\frac{1}{T_2} - \frac{1}{373 \text{ K}} \right) \exp \left(\frac{h_{\text{vap}}}{R} \right)$

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(a) The Clausius-Clapeyron equation: $dP_i \text{ sat } P_i \text{ sat} \approx \frac{h_i \text{ vap}}{RT^2} \text{ or } \ln$

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you should be able to: „Explain why it is
convenient to use the thermodynamic

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property Gibbs energy to determine pure species phase equilibrium. Discuss the balance between energetic and entropic effects at equilibrium.

Chapter 6 Koretsky | Chemical Equilibrium | Phase (Matter)

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An approximate solution can be found if
we combine Equations 1.4 and 1.5:
molecular $keV_m \approx 1 \times 10^{-22}$
molecular $keT \approx 3 \times 10^{-23} m kTV$
 $3 \times 10^{-23} \times 10^5 \approx 3 \times 10^{-18}$ Assume the
temperature is 22 °C.

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